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Experiment 1: Determine the Relationship Between the Volume and Pressure of Propane gas Lab Results 1. Record the pairs of data for pressure and propane volume in the table below.

Volume of Gas (mL) Pressure (atm) 150 1.00 130 1.15 110 1.36
90 1.67 70 2.14 50 3.00 Data Analysis 2. Create and save a graph of pressure in atm (y axis) versus ...

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Calculate the final volume of the reaction mixture after the contents of beaker B are added to beaker A. Report your answer in liters. All $\text{S}_2\text{O}_3^{2-}$ is consumed at the end of the reaction. Therefore, the moles of $\text{S}_2\text{O}_3^{2-}$ consumed can be calculated using the equation below. (moles $\text{S}_2\text{O}_3^{2-}$) consumed = $M_{\text{stock}} \times (V_{\text{stock}})$. The I_3^- produced in reaction 1 reacts with $\text{S}_2\text{O}_3^{2-}$ in reaction 2 as ...

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Please Help!!! Chemical Kinetics Lab Results Recor ...

Chemistry. A certain amount of methane is burned. Using stoichiometry it is calculated that 18.85 g of carbon dioxide should be produced. Carried out in the lab, 16.34 g of carbon dioxide are actually produced. What is the theoretical yield . asked by Abby Walker on July 8, 2012; Chemistry. I am so confused with my chemistry lab right now.

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